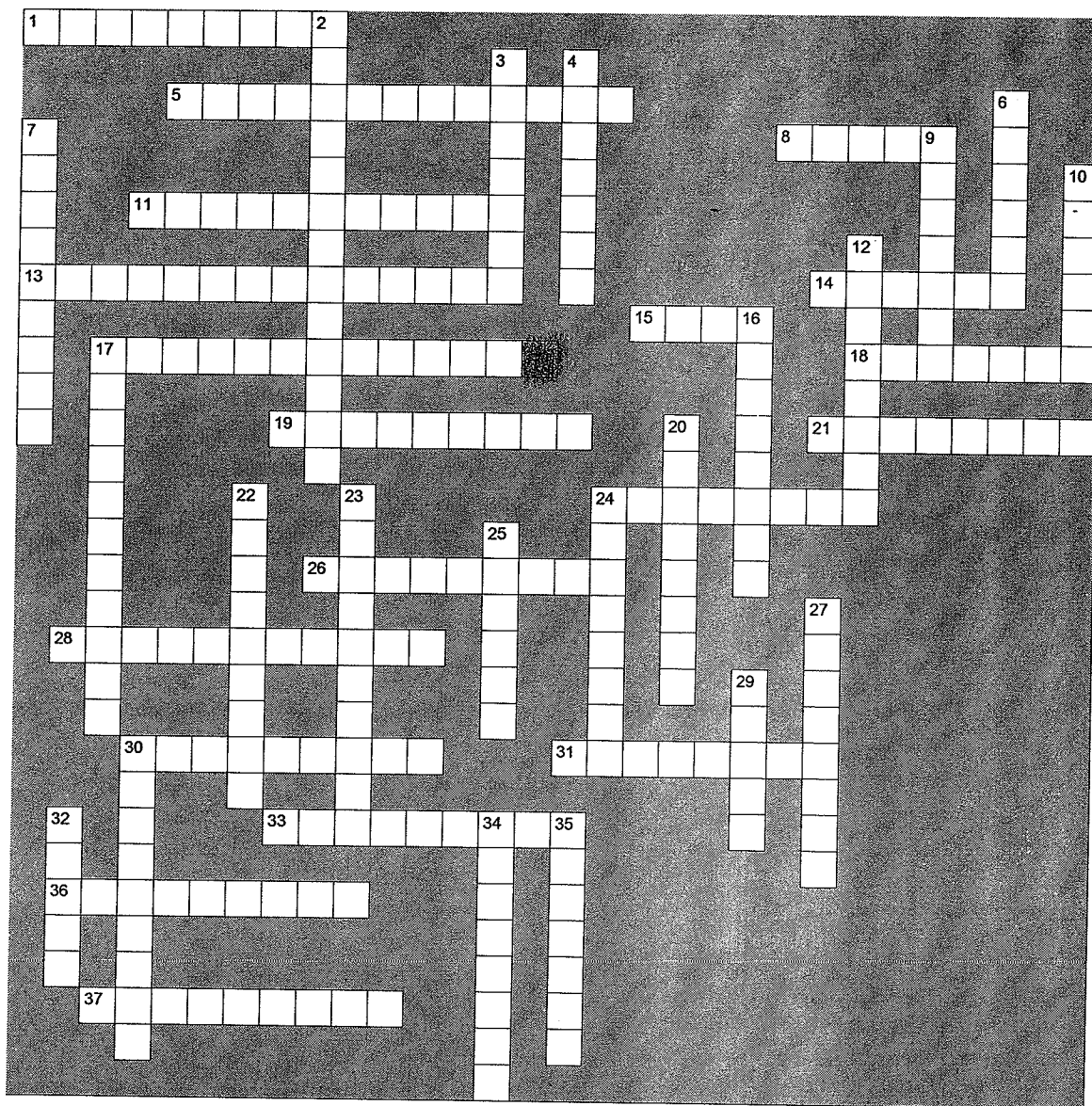


Final Exam Review Ch 6,7,8,9,15



Across

- a solution that is holding the maximum amount of substance in solution
- process of using a balanced equation to calculate masses of reactants and products
- to change from grams to _____ you divide by the molar mass
- a solid that is produced when 2 solutions react
- amu stands for _____ (3 words)
- in a _____ replacement, 2 compounds react with each other
- avagadro's number of anything is known as a _____
- a solution in which a large amount of substance is dissolved
- the atomic masses on the periodic table are actually the _____ mass of all isotopes of an element
- formula of the actual molecules of a compound, may not be simplified ratio
- the lowering of a solution's concentration by adding water
- to find the number of atoms of an element, you must _____ moles by Avagadro's number
- formula which represents the lowest ratio of the elements in a compound
- the amount of substance which can be dissolved depends upon the _____
- number which tells the number of atoms of an element in a compound
- substance that is present before a reaction takes place
- in a _____ reaction, a single compound is the only product
- a process used to determine the concentration of a solution, often an acid
- to convert from one substance in an equation to another, you use the _____ (2 words)

Down

- in a _____ reaction, a single compound is the only reactant
- the _____ yield is the amount that is actually produced compared to how much should be made
- substance that is formed in a chemical reaction
- the substance that is dissolved in a solution
- 6.022×10^{23} is known as _____ number
- a substance that dissolves in water is said to be _____ in water
- in a _____ replacement, an element reacts with a compound
- moles of substance dissolved divided by liters of solution
- numbers, letters, and symbols that represent a chemical reaction
- the number in front of a chemical formula in an equation, used to balance the equation
- equations must be _____ because of the law of conservation of mass
- the mass of one mole of an element or compound
- _____ reactions involve a hydrocarbon reacting with oxygen to produce water and carbon dioxide
- the molecular formula is a whole number _____ of the empirical formula
- a solution in which a small amount of substance is dissolved
- the _____ reactant controls the amount of product that can be made
- to change from moles to _____ you multiply by the molar mass
- these ions are removed when writing the net ionic equation
- the state symbol (aq) represents a substance dissolved in _____
- a homogeneous mixture
- the substance that is doing the dissolving in a solution